

Name _____

Period _____

Oxidation Number Rules:

1. A pure element has an oxidation number of 0.
2. The oxidation number of an element in a monatomic ion equals the charge of the ion.
3. The sum of the oxidation number of all the elements in a compound equals 0.
4. The sum of the oxidation numbers of all the elements in a polyatomic ion equals the charge on the ion.

Directions: In the following questions, give the oxidation number of the indicated atoms/ion.

1. N in N_2O_3

6. Mn in MnO_2

2. S in H_2SO_4

7. S in SO_3^{2-} (Note: Overall charge = -2)

3. C

8. Mg^{2+}

4. Na in NaCl

9. Ni in $\text{Ni}(\text{ClO}_2)_2$

5. S in Al_2S_3

10. S in H_2S